

MOKASA JOINT EVALUATION EXAMINATION

Kenya Certificate of Secondary Education

233/3

CHEMISTRY

PAPER 3

PRACTICAL

MARCH/APRIL 2011

CONFIDENTIAL

In addition to the apparatus found in chemistry laboratory, each candidate requires

1. 100cm³ of 0.02M potassium permanganate solution.
2. 100 cm³ of acidified potassium iodide solution.
3. 100cm³ of sodium thiosulphate
4. Burette
5. Pipette
6. Glass rod
7. 50cm³ measuring cylinder.
8. 100 cm³ measuring cylinder.
9. Spatula
10. About 2g solid P in a boiling tube.
11. 5 test tubes in a test tube rack
12. Boiling tube

Access to:

- 2M NaOH
- Ammonia solution
- Barium chloride (acidified)
- Lime water – calcium hydroxide
- Source of heat
- Acidified potassium permanganate
- Bromine water.

Notes

1. Potassium iodide is prepared by dissolving 1g of potassium iodide in about 20cm³ of distilled water in a conical flask and acidified with an equal volume of dilute sulphuric or hydrochloric acid.
2. Starch indicator is prepared by mixing 10g of starch to a thin paste with water in an evaporating dish, then pouring the paste into about 250cm³ of boiling water.
3. Solid P is a mixture of zinc carbonate and sodium sulphate in equal proportions.
4. Solid X is potassium chlorate.