

233/3
CHEMISTRY
PRACTICAL
PAPER 3
JULY / AUGUST 2010
CONFIDENTIAL

BORABU – MASABA NORTH DISTRICTS JOINT EVALUATION TEST - 2010
Kenya Certificate of Secondary Education (K.C.S.E)

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The information contained in this paper is to enable the head of the school and the teacher in charge of chemistry to make adequate preparation for this year's chemistry practical examination. NO ONE ELSE should have access to this paper or acquire knowledge of its contents. Great care MUST be taken to ensure that the information herein does not reach the candidates- either directly or indirectly.

The teacher in charge of chemistry should NOT perform any of the experiments in the same room as the candidates nor make the results of the experiments available to the candidates or give any other information related to the experiments to the candidates. Doing so will constitute an examination irregularity which is punishable.

In addition to the fittings and apparatus found in a chemistry laboratory, each candidate will require the following:

- A.
1. 100cm³ of solution P
 2. 120cm³ of solution R
 3. 75cm³ of solution T
 4. One 250ml volumetric flask
 5. One pipette, 25.0ml and a pipette filler
 6. One burette 0 – 50ml
 7. 2 Labels
 8. 2 conical flasks
 9. One thermometer (-10⁰C – 110⁰C)
 10. One 10ml measuring cylinder
 11. One 100ml beaker
 12. 0.2g of Sodium hydrogen carbonate
 13. 400cm³ of distilled water
 14. 10cm³ of solution H in a boiling tube
 15. 0.5g of solid J
 16. A clean metallic spatula
 17. One boiling tube
 18. Test-tube holder
 19. 6 test-tubes
 20. One filter funnel
 21. One filter paper whatman 125mm No.1

Access to

- B.
1. Bunsen burner
 2. Bromine water supplied with a dropper
 3. Methyl orange supplied with a dropper
 4. Acidified potassium dichromate (VI)
 5. 2M ammonia supplied with a dropper
 6. 2M Sodium hydroxide supplied with dropper
 7. 0.5M Barium chloride
 8. 2M hydrochloric acid.

NOTES:

1. Solution P is prepared by dissolving 79.30g of (oxalic) ethanedioic acid in 600cm³ of distilled water and diluting it to one litre solution.
2. Solution R is prepared by dissolving 5.56g of potassium carbonate in 700cm³ of distilled water and diluting it to one litre solution.
3. Solution T is prepared by dissolving 40g of sodium hydroxide pellets in 600cm³ of distilled water and diluting it to one litre of solution
4. Acidified potassium dichromate (VI) is prepared by dissolving 25g of solid potassium chromate (VI) in about 600cm³ of 2M sulphuric acid and diluting to one litre of solution.
5. Bromine water is prepared by diluting 1cm³ of liquid bromine with 100cm³ of distilled water in fume cup board.
6. Solution H is prepared by adding 75g hydrated copper (II) sulphate and 15g of Zinc nitrate into 500cm³ of distilled water and diluting it to one litre solution.
7. Solid J is malleic acid.