

CONFIDENTIALS

CHEMISTRY PRACTICAL JUNE 2016 KASSU JET

INSTRUCTIONS TO SCHOOLS

In addition to usual provisions and fitting in the science laboratory each candidate is expected to have the following.

- One pipette
- One burette
- One pipette filler
- About 100ml of solution M
- About 100ml of solution N
- Two 250ml conical flask
- One boiling tube
- 10ml measuring cylinders
- A stop watch
- 50ml solution W_1
- 500ml distilled water provided in wash bottle
- 100ml of 1M dilute hydrochloric acid (W_{II})
- One plane paper
- One 100ml glass beaker
- About 2.0g of solid P
- Metallic spatula
- Two filter papers
- One filter funnel
- Seven dry clean test tubes
- One dry boiling tube
- Acidified potassium chromate vi solution
- Bromine water
- Solid F (maleaic acid about 2g)

Access to:

- Source or means of heating
- 2M NaOH supplied with a dropper
- 2M NH_4OH supplied with a dropper
- 0.2M lead nitrate solution K supplied with a dropper
- 2M sulphuric acid supplied with a dropper
- Dilute nitric acid
- Universal indicator solution

Notes

- Solution M is made by accurately weighing 3.95g of potassium Manganate (VII) and dissolving it in 400ml of 1M H_2SO_4 and making up to solution 10, 1litre mark.
- Solution N is prepared by dissolving 49.0g ammonium ferrous sulphate, $(NH_4)_2SO_4 \cdot FeSO_4 \cdot 6H_2O$ and dissolving it in 400ml distilled water and making it up to 1litre mark.
- Solid P is a mixture of zinc chloride and magnesium carbonate in the ratio of 1;2
- Solution K is 0.2M lead nitrate solution.
- Substance W_1 is prepared by dissolving 40.0g sodium Thiosulphate in 500ml distilled water and making it up to 1.0 litre of solution
- Solution (W_{11}) 1M HCl is prepared by dissolving 86ml conc HCl in 600ml distilled water and making solution to 1litre mark.